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Chem 111 Chemical Equilibrium Worksheet Answer Keys. WORKSHEET: CHEMICAL EQUILIBRIUM Name Last Ans: First FOR ALL EQUILIBRIUM PROBLEMS, YOU MUST: 1) Write all equilibrium equations 2) Write all equilibrium concentrations 3) Write all equilibrium expressions SET A: a) What is the equilibrium Constant expression for the reaction:  $3 \text{Fe}(s) + 4 \text{H}_2\text{O}(g) \rightleftharpoons 3 \text{FeO}(s) + 4 \text{H}_2(g)$  b) The equilibrium constant, Kc, for the reaction:

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Figure 9.5.1: Equilibrium in reaction:  $\text{H}_2(g) + \text{I}_2(g) \rightleftharpoons 2 \text{HI}(g)$ . Chemical equilibrium can be attained whether the reaction begins with all reactants and no products, all products and no reactants, or some of both. The figure below shows changes in concentration of  $\text{H}_2$ ,  $\text{I}_2$ , and  $\text{HI}$  for two different reactions.

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Answer. The given reaction is  $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$ ,  $K_c = 4.08 \times 10^{-4}$ . Reversing the reaction gives the proper reactants and products for the target reaction, but with the wrong stoichiometry. Reversing the reaction also means that the new equilibrium constant is the inverse of the original equilibrium constant.

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Where To Download Chapter 18 Review Chemical Equilibrium Answers REVIEW Chemical Equilibrium SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. Match the solution type on the right to the corresponding relationship between the ion product and the  $K_{sp}$  for that solution, listed on the left. \_\_\_\_\_ The ion product ...

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equilibrium expression.  $K = \frac{(\text{C})^l (\text{D})^m}{(\text{A})^j (\text{B})^k}$ . equilibrium constant.  $K$ ; only changes when the temperature changes! ; depends on the ratio of the concentrations ; NEVER include pure solids or pure liquids ; one  $K$  at a specific temperature. equilibrium position.

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The equilibrium mixture contained  $1.37 \times 10^{-2} \text{ M HI}$ ,  $6.47 \times 10^{-3} \text{ M H}_2$ , and  $5.94 \times 10^{-4} \text{ M I}_2$ . Calculate  $K$  and  $K_p$  for this reaction. Answer:  $K = 48.8$ ;  $K_p = 48.8$

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Chemical Equilibrium Review Answer Key True or false, chemical equilibrium is a state in which the forward and reverse reactions take place at different rates. nope! The eq position of a reaction is given by the relative \_\_\_\_\_ of the systems components at the

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View Answer. A reversible reaction at equilibrium can counteract the stress of an increase in pressure by: a) favoring the formation of more gaseous products b) favoring the formation of fewer ...

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If all are equal, answer E. a. the concentrations of reactant and products are equal b. the rate constants for the forward and reverse reactions are equal c. the time that a particular atom or molecule spends as a reactant and product are equal d. the rate of the forward and reverse reaction e. all of the above are equal.

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increases. Chemical equilibrium is when the rate of the 'forward reaction equals the rate of the reverse reaction. 13. In closed container no external affect the reaction, so concentrations can become constant. 14. No. Equilibrium implies that the forward and reverse reactions occur at the same rate. 15. The forward and reverse reactions are

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