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ways in which electrons are

arranged in various orbitals around

the nuclei of atoms are called

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electron configurations. Three rules—the aufbau principle, the Pauli exclusion principle, and Hund ' s Page 1/5.

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the modern description of the

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electrons in atoms was proposed that electrons travel in circular orbits around the nucleus Bohr's contribution to the development of atomic Structure: (A. was referred to as the "plum pudding model," B. was the discovery that electrons surround a dense

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nucleus, C. was proposed that electrons travel in circular orbits around the nucleus, D. is the quantum mechanical model)

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ways in which electrons are arranged in various orbitals around the nuclei of atoms are called electron configurations. Three rules—the aufbau principle, the Pauli exclusion principle, and

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Hund's Page 1/5. Electrons In

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I can place a maximum of 2

electrons in the first orbital and 4

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in the 2nd orbital, which can hold up to 8 electrons. I don't need 8 electrons in the 2nd orbital so I am done because $2 + 4 = 6$ and all the electrons have been placed in orbitals. 5. Draw the atomic model for carbon with the protons, neutrons and electrons and orbitals

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~~Practice Problems Worksheet...~~

Valence Electrons, Ions, and the
Periodic Table. Valence Electrons.

- Valence electrons are electrons in the outer energy level of an atom. Practice!
- Find the number of valence electrons for: 1.Li 2.P 3.I 4.Ar 5.Ca 6.O 7.He. Check It!

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1.Li = 1 2.P = 5 3.I = 7 4.Ar = 8

5.Ca = 2 6.O = 6 7.. He = 2 (not 8)!

~~Valence Electrons, Ions, and the Periodic Table~~

protons of an atom for a specific element cannot change. of protons

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in an atom cannot change. In actuality, the number of protons is what identifies an element.

Electrons, because they are in the electron cloud, can be removed or added depending on the properties of the element. Be sure to record any extra helpful notes, vocabulary

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terms, and practice sample
problems.

~~This movement of electrons
between atoms is called ...~~

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Atoms. Matter is anything that
takes up space and has mass. All
matter is made of atoms. Atoms
are the basic building blocks of

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matter. They make up everything around us; Your desk, the board, your body, everything is made of atoms! Atoms are too small to see without powerful microscopes.

~~Atoms - Coventry Schools~~

The atom is a ball of positive

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charge with electrons stuck into the ball. Electrons are arranged in concentric circular paths around the nucleus. Most of an atom ' s mass is concentrated in the small, positively charged nucleus. The electrons surround the nucleus and the rest of the atom is empty

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space. false farther It is called a
quantum.

~~SECTION 5.1 MODELS OF THE
ATOM (pages 127–132)~~

In chemistry and atomic physics,
an electron shell may be thought of
as an orbit followed by electrons

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around an atom's nucleus. The closest shell to the nucleus is called the "1 shell" (also called the "K shell"), followed by the "2 shell" (or "L shell"), then the "3 shell" (or "M shell"... ..

~~Chapter 5 Electrons In Atoms~~

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A single covalent bond involves

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one shared pair of electrons. In many compounds, atoms will share electrons to enable their valence shell to become like the nearest Noble Gas. This is normally 8 electrons (the “ Octet Rule ”), apart from Hydrogen. There are exceptions (see next section).

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~~Chemical Bonding Worksheet~~

It aligns to the NGSS Practice of the Scientist of Developing and using models because while students cannot physically manipulate valence electrons at the atomic level they can model

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how atoms gain a full valence shell through gaining or losing electrons. It aligns to the NGSS Crosscutting Concept of Structure and Function.

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Guided Notes An atom is the smallest part of an _element_ which can take part in a chemical reaction. The atom

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