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In this lab, we are given an unknown carbonate and our job is to find out what it is. The pre-lab work gives the basic process on how to determine after all the measuring in the actual

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Unknown 11
experiment is done. ... The unknown solids are slightly toxic by ingestion and are skin irritants. Handle the crucible only with tongs. ... 11.
Another 400 mL ...

Gravimetric Analysis of an Unknown Carbonate - A. Sedano ...

Introduction viii PS-2877PS-2877
inquiry possibilities for students ' investigations see the suggestions in Using these Labs with the AP and the IBO Programs in this Introduction. Additionally, this manual presents teacher-developed laboratory activities using 21 st-century technologies to help you and your students explore topics, develop scientific inquiry skills, and

Advanced Chemistry Teacher Guide
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titration, using 10.0 mL of 1.5 M HCl to determine the molarity of a solution of NaOH with an unknown concentration with the use of the indicator phenolphthalein. ... In conclusion, it takes on average 11.9 mL of the NaOH of unknown concentration to neutralize the 1.5 M HCl. This is slightly more than the ...

Titration Lab - AP Chemistry
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relative amounts of reactants and
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templates. the equilibrium
position refers to the relative amounts
of reactants and products in the
system at the point of equilibrium. 1:
Introduction to Equilibrium. mL/s
(many other possible answer would
work) 4.

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Nothing in the history of mass fatality
events prepared America ' s forensic
community for the task of identifying
those who died when terrorists
attacked the World Trade Center in
New York City on September 11,
2001. The number of victims, the

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condition of their remains, and the duration of the recovery effort made their identification the greatest forensic challenge ever undertaken in this ...

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This lab is adapted from: Kristen L. Cacciatore and Hannah Sevia, Teaching Lab Report Writing through Inquiry: A Green Chemistry Stoichiometry Experiment for General Chemistry, Journal of Chemical Education, 83(7), 1039, 2006. 4-3
Therefore, the theoretical atom economy is: $100\% \times \frac{243.8}{111.7} = 45.82\%$.
100% Mass of all reactants

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