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Aqueous Solution

An aqueous solution is a solution in which
the solvent is water, whereas in a

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nonaqueous solution, the solvent is a substance other than water. Familiar examples of nonaqueous solvents are ethyl acetate, used in nail polish removers, and turpentine, used to clean paint brushes. In this chapter, we focus on reactions that occur in aqueous solution.

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You have seen that mixing solutions of two ionic compounds can sometimes result in the formation of an insoluble salt called a precipitate. □Some combinations of solutions produce precipitates, while others do not.

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ANSWERS
VIDEO 8: In this video we introduce chemical reactions, focusing on aqueous solutions and precipitation reactions.

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But 113 Reactions In Aqueous Solution
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of aqueous solutions of silver nitrate with
sodium chlo-ride to form solid silver
chloride and aqueous sodium nitrate is a
double-replacement reaction. The reaction

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Several types of reactions occur in water. When water is the solvent for a reaction, the reaction is said to occur in aqueous solution, which is denoted by the

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abbreviation (aq) following the name of a chemical species in a reaction. Three important types of reactions in water are precipitation, acid-base, and oxidation-reduction reactions.

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Calculations for Acid -Base Titrations

Oxidation-Reduction Reactions. 4.

Balancing Redox Equations 5. Adding in
H⁺, OH⁻, or H₂O to Balance

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Answers

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Skin and eye irritant. Zinc nitrate :
Irritating to eyes, skin, respiratory system,
and digestive system. Magnesium:

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flammable; contact with water releases flammable gas. May cause skin and eye irritation. Seek medical attention if irritation occurs. Do not induce vomiting. Phosphoric acid : Danger, corrosive.

CHM-113L-L5-ChemicalReactionsInAqueousSolutions ...

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aqueous solutions . are mixed, and then test your predictions in the laboratory. During the previous discussion period, your lab instructor lectured on the topic of reactions in aqueous solution with examples of the correct way to write a molecular equation, an ionic equation, and the overall net ionic equation for several

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An aqueous solution is any solution in which water (H_2O) is the solvent. In a chemical equation, the symbol (aq) follows a species name to indicate that it is in aqueous solution. For example, dissolving salt in water has the chemical reaction:

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Aqueous Solution Definition in Chemistry

The reaction of aqueous solutions of silver nitrate with sodium chloride to form solid silver chloride and aqueous sodium nitrate is a double-replacement reaction. The reaction is shown in Figure 11.11. $\text{AgNO}_3(\text{aq}) + \text{NaCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{NaNO}_3(\text{aq})$ This is the way you have been writing

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equations involving aqueous solutions of ionic compounds.

11.3 Reactions in Aqueous Solution

Substituent Effects on an Inverse Electron

Demand Hetero Diels-Alder Reaction in

Aqueous Solution and Organic Solvents:

Cycloaddition of Substituted Styrenes to

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Di(2-pyridyl)-1,2,4,5-tetrazine. The
Journal of Organic Chemistry 1996 , 61
(6) , 2001-2005.

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